

Name

Student's No.

Delegation

**Student's Report for Practical
Problem I**

1. Identification of unknown solutions

labels of solutions	A	B	C	D	E
chemical formulae					

2. Equations for chemical reactions taking place in each experimental step

step one (with the first reagent you choose)

step two (with the second reagent you choose)

step three (with the third reagent you choose)

If you proceed with more steps, write the corresponding chemical equations continuously.

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**Student's Report for Practical
Problem II**

1. Equations for the main chemical reactions having taken place in the preparation of $\text{Cu}(\text{OH})_2$.

2. Mass of $\text{Cu}(\text{gly})_2 \cdot x\text{H}_2\text{O} =$ (g)

Percent yield

The expression for calculation

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**Student's Report for Practical
Problem III**

1. Standardization of Na₂S₂O₃ Solution

1) The two equations for chemical reactions taking place during the standardization of Na₂S₂O₃.

i

ii.

(2) Volumes of Na₂S₂O₃ Solution

$V_1 =$ cm^3 $V_2 =$ cm^3
 $V_3 =$ cm^3

$V(\text{mean value}) =$ cm^3

(3) Concentration of Na₂S₂O₃ = $\text{mol}\cdot\text{dm}^{-3}$

Expression for the calculation:

2. Determination of Cu(II) in Cu(gly)₂·xH₂O

(1) Chemical equation for the reaction between Cu²⁺ and I⁻

(2) Mass of Cu(gly)₂·xH₂O = g

3) Volumes of Na₂S₂O₃ solution

$V_1 =$ cm^3 $V_2 =$ cm^3 $V_3 =$ cm^3

$V(\text{mean value}) = \quad \text{cm}^3$

**(4) Mass % of Cu(II) in $\text{Cu}(\text{gly})_2 \cdot x\text{H}_2\text{O} = \quad \%$
Expression for the calculation:**

(5) X Value in $\text{Cu}(\text{gly})_2 \cdot x\text{H}_2\text{O}$

$X = \quad$ (with precision of 0.01)

Expression for the calculation: